ECAT Chemistry Chapter 4 Liquids & Solids

Sr	Questions	Answers Choice
1	Escape of high energy molecules from the surface of a liquid is called	A. Sublimation B. Distillation C. Condensation D. Evaporation
2	Which one of the following is the weakest intermolecular force	A. Dipole induced dipole forces B. Ionic dipole forces C. Electrostatics forces between ions D. Dipole-dipole forces
3	H_2S is a gas while H_2O is a liquid at room temperature. It is due to:	A. Less inter-molecular forces in water. B. Covalent bonding H-O in water molecule. C. Hydrogen bonding in water molecules D. Ionic character in water molecules
4	Hydrogen bonding is present in one of the following pairs:	A. NH₃ B. H₂O C. HF D. All of above
5	Amorphous solids	A. Have sharp melting point B. Undergo clean cleavage when cut with knife C. Have perfect arrangement of atoms D. Can possesses small regions of orderly arrangement of atoms
6	London dispersion forces are also called	A. Hydrogen bonding B. Debye forces C. Van de Waal's forces D. Instantaneous dipole-induced dipole forces
7	The density of water decreases, When it is freezed at 0°C	A. Change of bond length B. Change of bond angles C. Cubic structure of ice D. Empty spaces present in the structure of ice
8	HF has exceptionally low acidic strength due to:	A. Smaller size of fluorine. B. Stronger polar bond between H and F. C. Electronegativity of fluorine D. Strong hydrogen bonding
9	Polypeptide chains are coiled about one another into a spiral by	A. Ionic bonds B. Covalent bonds C. Van der Waal's forces D. Hydrogen bonds
10	The strongest forces are:	A. Debye froces B. London dispersion C. Dipole-dipole attraction D. Hydrogen bonding
11	Amorphous solids:	A. Have a sharp melting point B. Undergo clean cleavage when cut with knife C. Have a perfect arrangement of atoms D. Can possesses small regions of orderly arrangement of atoms
12	Diamond is a bad conductor because:	A. Ithas tight structure B. It has a high density C. There is no free electron present in the crystal of diamond of conduct electricity is transparent to light
		A. increasing the pressure only B. Lowering temperature and

13	The gases can be converted into liquids by	increasing pressure C. Increasing pressure and bringing temperature below critical point D. Lowering temperature only
14	Trend of boiling point of halogens from fluorine to Iodine is that it	A. Decreases B. Is negligible C. Increases D. Remains constant
15	A liquid on evaporation causes:	A. Heating effect. B. Cooling effect. C. Suffocation . D. All of above
16	Diamond is a bad conductor because	A. It has a tight structure B. It has a high density C. there is no free electron present in the crystal of diamond of conduct electicity D. Is transparent to light
17	Which of the following liquids has low vapor pressure at 25°C:	A. Diethyl ether B. Acetone. C. Water. D. Ethyl alcohol.
18	Rate of evaporation and rate of condensation at equilibrium	A. Become very low B. Become very high C. Become equal D. Can never be equal
19	Ionic solids are characterized by :	A. Low meeting points B. Good conductivity in solid state C. High vapor pressure D. Solubility in polar solvents
20	The only forces are London dispersion forces among the	A. Atoms of He in gaseous state at high temperature B. Molecules of water in liquid state C. Molecules of solid I₂ D. Molecular of hydrochloric acid gas