ECAT Chemistry Chapter 7 Thermo Chemistry

Sr	Questions	Answers Choice
1	When a system absorbs energy, the sign of delta E is	A. Neither positive nor negative B. Negative C. Positive D. None of above
2	A reaction will also be called a spontaneous if	A. It does not need energy to start with B. It needs energy to carry the whole process C. It needs energy at the end of reaction D. It needs energy to start with
3	The subject matter of first law of thermochemistry is based on	A. First law of Thermochemistry B. First law of Thermodynamics C. Second law of Thermochemistry D. Second law of Thermodynamics
4	The energy units in which heat changes usually expressed in SI-system are:	A. Joule. B. Calorie. C. Kilo Joule. D. Both (a) and (c)
5	If an endothermic reaction is allowed to take place very rapidly in the air, the temperature of the surrounding air	A. remains constant B. increases C. decreases D. remain unchanged
6	Which of the following statements is contrary to the first law of thermodynamics?	A. Energy can neither be created no destroyed. B. One form of energy can be transferred into an equivalent amount of the kinds of energy. C. In a adiabatic process, the work done is independent of its path. D. Continuous production of mechanical work without supplying an equivalent amount of heat is possible.
7	The study of heat changes accompanying a chemical reaction is known as	A. Thermochemistry B. Biochemistry C. Physical chemistry D. Analytical chemistry
8	The net heat change in a chemical reaction is same, whether it is brought about in two or more different ways in one or several steps. It is known as:	A. Henry's law. B. Joule's principle. C. Hess's law. D. Law of conservation of energy.
9	When a piece of zinc is added to the copper sulphate solution, _____ colour of solution disappear	A. Pink B. Purple C. Blue D. Brown
10	A process which takes place on its own without any outside assistance and moves from a non-equilibrium stat towards an equilibrium stat is termed as:	A. Spontaneous process. B. Natural process. C. Non-Spontaneous process. D. Both (a) and (b).
11	In endothermic reactions, the heat contents of the:	A. Products equal to the reactants. B. Reactants more than that of products. C. Products more than that of reactants D. Both (b) and (c)
12	Calorie is equivalent to	A. 0.4184 J B. 41.84 J C. 4.184 J D. 418.4 J
13	The change in heat energy of a chemical reaction at constant temperature and pressure is called	A. entahlpy change B. heat of sublimation C. bond energy

		D. Internal energy change
14	The majority of reactions which give stable products are:	A. Exothermic. B. Isothermal. C. Endothermic. D. Both (b) and (c).
15	The net heat in a chemical reaction is same, whether it is brought about in two or more different ways in one or several steps. It is known as	A. Henry's law B. Joule's princile C. Hess's law D. law of conservation of energy
16	Which one of the following is not related to spontaneous process	A. Unidirectional B. Real C. Irreversible D. Artificial
17	Work is a	A. State function B. Only function C. Non-state function D. State
18	The subject matter of first law of thermochemistry is based on:	A. First law of Themochemistry. B. First law of Thermodynamics. C. Second law of Themochemistry. D. Second law of Thermodynamics.
19	Reaction between Zn and $CuSO_4$can be called a system under	A. Surrounding B. Observation C. system D. None of above
20	In endothermic reactions, the heat content of the:	A. Products is more than that of reactants. B. Reactants is more than than to products. C. Both (a) and (b). D. Reactants and products are equal.