Acid bases and Salts

Sr	Questions	Answers Choice
1	What is teh standard unit for the mass/volume concentration.	A. mol/kg B. g/cm3 C. mg/l D. g/dm3
2	When natural gas burns in air, which component is the limiting reactant.	A. Both components B. Neither component is a limiting react C. Air D. natural gas
3	Concentratio of solution is a measure of the amount of solute dissolved in given amount of.	A. solution or solvent B. Only Solute C. Solid crystals D. Catalyst
4	Glucose has an empirical formula of CH2O and a molecular mass of 180 .What is its molecular formula.	A. C2H4O3 B. C4H8O4 C. C6H12O6 D. C4H8O4
5	Benzene and acetyle both have the same emperical formula which is.	A. CH2 B. C2H C. CH D. CH4
6	To find the molecular formula from an empricial formula, the multiplier calculated by.	A. Emperical mass/molecular mass B. Molecular mass/Empirical formula mass C. Atomic mass x moles D. Actual yield x theoretical yield
7	Which of the following is an empirical formula of benzene.	A. C6H6 B. C6H12O6 C. CH D. 6CH
8	Empirical formula of magnesium chloride is.	A. MgCl2 B. 2 MgCl C. Mg2Cl2 D. 2 MgCl2
9	Emperical data required for titration calculations includes.	A. Volumes of two solutions used B. The atmospheric pressure only C. The weight of the burette D. None of these
10	It is esential for optimizing chemical reactions in many industries.	A. Stoichimetry B. Biochemistry C. Organic chemistry D. Electrochemistry
11	Volume occupied by 15 moles of ammonia gas at RTP.	A. 224 dm3 B. 160 dm3 C. 360 dm3 D. 265 dm3
12	Titration is a technique used to find the concentration of a solution by reacting it with.	A. Water only B. Solution of known concentration C. Insoluble salt D. Gaseous catalyst
13	what will the molar concentration of potassium hydroxide solution if 2.0 g it are dissolved in 100 cm3 of solution.	A. 0.36 mol/dm3 B. 053 mol/dm3 C. 0.70 mol/dm3 D. 0.45 mol/dme
14	How many moles of natural gas CH4 will be present in a cylinder if its volume is 18 dm3 at RTP?	A. 0.75 Mol B. 0.67 mol C. 0.85 mol D. 0.56 mol
15	the percentage mass of each element present in a compound is know as.	A. Molecular mass B. Percentage composition C. empirical ratio

		D. Molar volume
16	To determine the concentration of a solution via titration, you must know the	A. Density of the flask B. Room humidity C. Colours of the solute D. Balanced chemical equaiton
17	For titration purposes, you need to know the concentration of.	A. Both solutions beforehand B. The other solution C. Indicator D. Both a and b
18	The reactant that is consumed earlier in a reaction and controls the amount product formed is the.	A. Excess reactant B. Catalyst C. Limiting reactant D. Theoretical reactant
19	the pressure value used for standard RTP conditons is.	A. 100 kpa B. 1 atmosphere (760 torr) C. 273 torr D. 24 atmosphere
20	The quantity of product actually obtained when a chemical raetion is performed is called.	A. theoreticla yield B. Actual yield C. Percent purity D. Stoichiometric yield