Acid bases and Salts

Sr	Questions	Answers Choice
1	What is teh room temperature for standard conditions.	A. O oC B. 25 oC C. 75 oC D. 100 oC
2	what will the molar concentration of potassium hydroxide solution if 2.0 g it are dissolved in 100 cm3 of solution.	A. 0.36 mol/dm3 B. 053 mol/dm3 C. 0.70 mol/dm3 D. 0.45 mol/dme
3	Emperical data required for titration calculations includes.	A. Volumes of two solutions used B. The atmospheric pressure only C. The weight of the burette D. None of these
4	The reactant that is consumed earlier in a reaction and controls the amount product formed is the.	A. Excess reactant B. Catalyst C. Limiting reactant D. Theoretical reactant
5	the pressure value used for standard RTP conditons is.	A. 100 kpa B. 1 atmosphere (760 torr) C. 273 torr D. 24 atmosphere
6	To determine the concentration of a solution via titration, you must know the	A. Density of the flask B. Room humidity C. Colours of the solute D. Balanced chemical equaiton
7	At RTP, one mole of any gas occupies a volume of.	A. 22 dm3 B. 23 dm3 C. 24 dm3 D. 25 dm3
8	To find the molecular formula from an empricial formula, the multiplier calculated by.	A. Emperical mass/molecular mass B. Molecular mass/Empirical formula mass C. Atomic mass x moles D. Actual yield x theoretical yield
9	When natural gas burns in air, which component is the limiting reactant.	A. Both components B. Neither component is a limiting react C. Air D. natural gas
10	Benzene and acetyle both have the same emperical formula which is.	A. CH2 B. C2H C. CH D. CH4
11	The formula that represents the simpleast whole numebr ratio of atoms in compund is the	A. Empricial formula B. Molecular formula C. Molar volume D. Percentage composition
12	The quantity of product actually obtained when a chemical raetion is performed is called.	A. theoreticla yield B. Actual yield C. Percent purity D. Stoichiometric yield
13	Which factor does NOT chagne teh volume occupied by one mole of a gas a RTP.	A. Temperature B. Pressrue C. Size of the molecules D. Numebr of moles
14	Volume occupied by 15 moles of ammonia gas at RTP.	A. 224 dm3 B. 160 dm3 C. 360 dm3 D. 265 dm3
15	How many moles of HCl are needed to react with 1 mole of Ca CO3 in the given reaction.	A. 1 mole B. 2 moles C. 3 moles

		D. 4 moles
16	Titration is a technique used to find the concentration of a solution by reacting it with.	A. Water only B. Solution of known concentration C. Insoluble salt D. Gaseous catalyst
17	Empirical formula of magnesium chloride is.	A. $MgCl_2$ B. 2 MgCl C. Mg_2Cl_2 D. 2 $MgCl_2$
18	When natural gas burns in air, which component is the limiting reactant.	A. Both components B. Neither component is a limiting react C. Air D. natural gas
19	For titration purposes, you need to know the concentration of.	A. Both solutions beforehand B. The other solution C. Indicator D. Both a and b
20	Glucose has an empirical formula of CH_2O and a molecular mass of 180 .What is its molecular formula.	A. $C_2H_4O_3$ B. $C_4H_8O_4$ C. $C_6H_{12}O_6$ D. $C_4H_8O_4$