Hydrocarbons

Sr	Questions	Answers Choice
1	According to collision theory, what are collisons that result in the formation of products called.	A. Random Collsions B. Effective collisions C. elastic collisions D. Frequent Collsions
2	The numebr of collision per unit volume of the reaction mixture is called.	A. Collision energy B. Activation energy C. Collision frequency D. Collision force
3	Which types of reactions releases heat energy into the surroundings.	A. Neutralization B. Exothermic C. Endothermic D. Decomposition
4	In many cases the rate of reaction is double when the temperatur eis increased by.	A. 10 oC B. 20 oC C. 30 oC D. 40 oC
5	What typically happens to the rate of many reactions for every 10 oC rise in temperature.	A. It doubles B. It halves C. It stays the same D. It triples
6	the minimum energy required to start a chemicla reactions known as.	A. thermal energy B. Kinetic energy C. Activation energy D. Potential energy
7	If the reactants posses an energy higher than the activation eneryg, then the reaction will be	A. Slow B. Fast C. Not affected D. Instantaneous
8	Which form of a solid reactant will react the fastest with a liquid.	A. Finely divided powdered B. A few medium sized lumps C. One large crystal D. A flast sheet of the solid.
9	What process do enzymes catalyze to ripen fruits.	A. Freezing the fruit tissue B. Conversin of starches into sugars C. Conversio fo sogars into starches D. Increasing the oxygen content
10	Which of the followiof a catalyst.ng explains the increase in rate of reaction in the presence	A. Catalyst provides extra eneryg to the reactant molecules. B. Catalyst provides an alternative pathway which lowers the activations energy. C. Catalyst increases teh collision frequency D. Catalyst decreeases the collision frequency
11	Which of the following statements is correct.	A. Collisions with energy equal to or greater than the activation energy lead to reaction. B. Collision frequency is not related to the reaction rate. C. All collisions lead to a reaction D. Collisions with energy less than the activation energy lead to the reaction.
12	Enzymes are best described as.	A. Products of spoilage B. Biological Catalyst C. Inorganic reactants D. temperature inhibitors.
13	When CaO reacts with water, what happens to the temperatute.	A. It remains at 0 oC B. It increases C. It decreases D. It fluctuates randomly

14	Which gas production rate can be monitored to determine the best time for fruit harvasting.	A. Ethene B. Nitrogen C. Hydrogen D. Carbon monoxide
15	How does increasing the concentration of reactants affect the reaction rate.	A. It increases teh rate B. It increaes activation energy C. It hase no effect on the rate D. It decrease the rate.
16	Dissolving ammonium chloride in water makes the container feel cold. This is example of a/an	A. Fast reaction B. Activation reaction C. Endothermic reaction D. Exothermic reaction
17	What happens to the concentration of products as a reaction moves toward completion.	A. It stays constant B. It decreases C. It increases D. It becomes zero
18	When a reaction proceeds ahead, how the concentrations of reactants and products change.	A. Concentration of reactants and that of the products decreases. B. Concentratration of reactants decreases and tha of the products increases C. Concentration of both reactants and products decreases. D. Concentration of both reactants and products increases.