Biochemistry

Sr	Questions	Answers Choice
1	Which force is responsible for bonding the oppositely charged ions together in a salt.	A. Nuclear force B. Electrostaitic force of attraction C. Gravitational force D. Megnetic force
2	Ionic compunds typically have high melting points of.....electrostatic of attraction are present between the ions.	A. Strong B. Very Strong C. Weak D. Very Weak
3	What colour chagne indicates that the sodium hydroxide has been neutralized.	A. Colourless to pink B. Pink to colurless C. Blue to red D. Red to colourless
4	Which of these salts is typically insoluble in water.	A. Sodium Carbonate B. Ammonium Nitrate C. Lead Sulphate D. Potassium chloride
5	Which base is not soluble in water.	A. KOH B. Mg(OH)2 C. NaOH D. Na2CO3
6	In which state doionic compunds NOT conduct electricity.	A. Molten state B. Solid state C. Liquid state D. Aqueous solution
7	the attractive force break down and let the ions move freely during process of.	A. Melting B. Boiling C. Evaporation D. Condensation
8	Ions are held together by storng forces of attractions in	A. Liquid state B. Solid state C. Gaseous state D. All of these
9	To prepare soluble potassium Nitrate which acid and base should be used in a titration method.	A. Hydrochloric acid and Potassium hydroxide B. Sulpuric acid and potassium carbonate C. Nitric acid and potassium hydroxide D. Nitric acid and sodium hydroxide.
10	Which type of insoluble base can be used be react with a soluble acid.	A. Metal oxide B. Metal carbonate C. Meal hydroxide D. All of these
11	in the reaction of NaOH +HCl.....NaCl = H2O What does NaCl represent.	A. Hydrochlroic acid B. Sodium hydroxide C. Sodium chloride D. Hydrogen oxide
12	In which state do ionic compunds Not conduct electricity.	A. Dissolved in water B. Molten state C. Solid state D. None of these
13	Ionic compunds serve as	A. Weak electroytes B. Bad electroytes C. good electrolytes D. Strong electrolytes
14	Under ordinary conditions of temperatures and pressure, in what state do compounds exist.	A. Liquid state B. Crystaline solids C. Gaseous state D. Plasma State
		A. condense

15	Ionic compounds serve as strong electroytes becasue they..... completely into ions when dissolved in water.	B. Freeze C. Evaporate D. Dissociate
16	In reaction of an acid with an insoluble base, what is the step taken to remvoe the excess insoluble base from the mixture.	A. Filtration B. Evaporation C. Crystallization D. Stirring
17	All carbonates are insoluble in water with the exception of carbonates of.	A. Sodium B. Potassium C. Ammonium D. All of these
18	How should pure salt crystals be dried once they are filtered.	A. By heating B. By leaving in a beaker C. Between the folds of filter paper D. None of these
19	Higher charges and smaller ion size lead to stronger attrractions and hence	A. Higher boiling points B. Lower melting points C. Lower boiling points D. Higher melting points
20	When an ionic compound dissolves in water the forces of attraction between its ions.	A. Become stronger B. Break down C. Remain the same D. Completely finish