Chemical Bonding

Sr	Questions	Answers Choice
1	Which form of carbon is used as a lubricant?	A. Coal B. Diamond C. Charcoal D. Graphite
2	Number of electronsin nitrogen molecule is.	A. 2 B. 4 C. 6 D. 8
3	Which ionic compound has the highest melting point.	A. RbCl B. KCl C. LiCl D. NaCl
4	The compund formed by oppoite charges are known as.	A. Metallic solids B. Ionic compounds C. Non-polar Covalent compound D. None of the above
5	Atoms achieve stability by attainign electonic configuration of.	A. Alkali metals B. Coinage metals C. Inert Gases D. Alkaline earth metals
6	Which group of the periodic table has the tendency to gain electrons.	A. Group -1 B. Group -17 C. Group-2 D. Group -18
7	Which molecule contains a single vovalent bond.	A. CH4 B. C2H4 C. C2H2 D. O2
8	The boiling point of water is	A. O ^oC B. 100 ^oC C. 35 ^oC D. 25 ^oC
9	Whcih compound contains both covalent and ionic bonds.	A. MgCl₂ B. PCl5 C. NH₄Cl D. CaO
10	How many covalent bonds does C_2H_2 molecule have.	A. Two B. Three C. Four D. Five
11	Whcih one of the following is the weakest force among the atoms.	A. Intermolecular force B. Ionic force C. Metallic force D. Covelent forces
12	Attaining two electrons in the valence shell is called.	A. Octet rule B. Duplet rule C. Triplet rule D. All of these
13	Silicon belongs to Grou IVA . It has ....electons in the valence shell	A. 2 B. 6 C. 3 D. 4
14	Whcih of the following is not true about the formationof Na_2S	A. Each sodium atom loses one electron B. Sodium forms cation C. Each sulphur atom gains one electron D. Sulphur form anion
15	Every atom has a natural tendency to accommodate electrons in its valence shell	A. 2 or 6 B. 2 or 4

		C. 2 or 8 D. 2 or 10
16	Water has high boiling point as compared a alcohol due to	A. Low density B. High surface tension C. Hydrogen bonding D. High vapour pressure
17	Keeping in view the intermolecular forces of attraction, indicate which compound has the highest boiling point	A. H₂S B. HF C. NH₃ D. H₂O
18	All the noble gases have their valence electrons.	A. Incomplete B. Partialy filled C. Completely filled D. None of the above
19	Why is H_2Oa liquid while H_2S is a gas?	A. Becuase in waer, the atomic size of oxygen in smaller than that of Sulphur B. Because water can easily freeze into ice C. Because water is a polar compound ad tere exists strong forces of attrction betwene its molecules D. Because H₂O molecule is lighter than H₂S
20	How many electrons are involved in the formation of signle covalent bond	A. One B. Two C. Three D. Four