Equilibria

Sr	Questions	Answers Choice
1	Predict which components of the amosphere react in the presence of lightening.	A. N2 and H2O B. O2 and H2O C. N2 and O2 D. CO2 and O2
2	Which of the following does not happpen, when a system is at equilibrium state.	A. Reaction continues to occur in both the directions B. Concentration of reactants and products stop changig C. Forward and reverse reactions stop D. Forward and reverse rates become equal
3	Such reaction which continue in both directins are called.	A. Dynamic B. Irreversible C. Reversible D. Non- reactive
4	The reaction in which the products can recombine to formreactants are called.	A. Reversible Reaction B. Irreversible reactions C. Decomposition reactions D. Addition reactions
5	In the beginning the rate of reverse reaction is.	A. Slow B. very fast C. Moderate D. Negligible
6	An Inorganic chemistry places one mle of PCl5 in container A and one mole of each Cl2 and PCl3 in container B. Both the containers were sealed and heated to the same temperture to reach the stte of equilibrium Guess about the composition of mixtures in both the containers.	A. Both the containers wil have zero concentration of its reactants. B. Both the containers wil hae the same composition of mixtures C. Container A will have more concentraion of PCl3 than B. D. Container A will have less concentraion of PCl3 than B.
7	What conditon Should be met for the reversible reactio to achieve the state of equilibrium.	A. The concentratio of all the reactants and the prodduct should become constant B. all the reactants should be converted into the product C. 50% of the reactant shuld be converted into prodcuts. D. One of the product should be removed from the reaction mixture.
8	The characteristics of reversible reactions are the following except.	A. Product nver recombine to form reactants B. They never complete C. They have a double arrow between reactants and products D. The proceed in both ways
9	The colour of hydrated copper (II) sulphate solid is.	A. Black B. Pink C. White D. Blue
10	Whcih is true about the equilibrium state?	A. The forard reaction stops B. Both forward and reverse reactions stop C. Both foward and reverse reactions continue at the same rate D. The reverse reaction stops
11	Which type of reactions speed up gradully?	A. Decomposition reaction B. Forward reaction C. Reverse reactions D. Irreversibel reactions
		A. Because gteh solubiliyt of the gas increases B. Because the gas is dissolved

12	Why the gas starts coming out when you open a can of fizzy drink.	under pressure hence it comes out when pressur eis decreased C. Becaue the gas is insoluble in water D. Because the solubility of hte gas decreases at high pressure.
13	CaO or lime is used extensive in steel, glass and paper industries. It is produced inan exothermic reversitble reaction by the decompositon of lie . Choose the conditions to produce maximum amount of lime.	A. Heating at high temperatur ein an open vessel B. Heating at high temperatur ein a closed vessel C. Cooling it in a closed vessel D. Colling it in an open vessel
14	If reversibe reaction useful for preparing compounds on large scale.	A. Yes B. No C. They are useful only when equilirbium lies far to the left side D. They are useful only when equilibrium lies far to the right side
15	At what temperature , rate of ammonia formation and decomposition is the highest.	A. 200 ^oC B. 300^oC C. 400 ^oC D. 500 ^oC
16	In chemical reaction, the substances that combine are called.	A. Masses B. Materials C. Products D. Reactants
17	What will happen if the rates of forward and reverse reactions are very high	A. The reaction will be practiclly irreversible B. The equilibrium point will reach very soon C. The equilibiurm point will reach very late D. The reaction will not attain the state of dynamic equilibrium
18	When system is at quilibrium state.	A. The rate of the forward and reverse rections become equal B. The concentrationof reatants and product becoes equal C. The oposing reactions stop D. The rate of the reverse reactio becomes very low
19	A complete reaction is in which	A. Only 10% reactants covert into products B. All the reactants covrt into products C. All the reactants do no covrt into products D. Half reactants covert into produts
20	Which compound is used a thinner is paint industry?	A. H2O B. C2H3OH C. CH3COOC2H5 D. CH3COOH