Electrochemistry

Sr	Questions	Answers Choice
1	Electroyte in electroplating of silver	A. AgNo3 B. NaCl C. CuSO4 D. HNO3
2	Electrons in a galvanic cell flow from	A. Anode to cathode B. Cathode to anode C. Salt bridge to cathode D. Elecrolyte to anode
3	A positive value for the standared electrode potential of a metal ion metal half cell indicate that	A. The metal is stong reducing agent B. The metal ion is redily oxidized C. The metal ion is redily reduced D. The metal will readily displace hydrogen from dilute acids.
4	The reducing agent in a redox reaction	A. Gains electrons B. Loss electrons C. Acceptw protons D. Donates protons
5	In an electrolysis exprement if a charge 96,500 Coluombs is passed trough a solution, the amount of substance liberated or deposited at the electrode is directly relatd to.	A. Mass numebr of the ion B. One mole of electrons being transferred C. Avogadro's nmber of ions being discharged D. Standrd electrode potentil of the metal iion
6	Electrolysis of CuSO4 using copper eletrodes results in	A. Increase in eletolyte concentration B. Decrease in Cu2+ concentration C. No change in electrolyte composition D. Formation of new compound
7	Which element has Eo = 0.00 V?	A. H+ B. H2 C. SHE D. All of the above
8	The reaction at the cathode is alwyas	A. Oxidation B. Neutral C. Reduction D. Ionization
9	Whcih metal will displace Cu from CuSO4 Solution.	A. Zn B. Ag C. Hg D. Pb
10	The oxidation number of sulfur in H2SO4 is	A. +2 B. +4 C. +6 D. +8
11	Which of the following metals would most raadily displace hydrogen gas from dilute acids.	A. Copper B. Silver C. Magnesium D. Platinum
12	Redox reactions always involve.	A. Gain of protons B. Transfer of electrons C. Loss ofneutrons D. Nuclear change
13	A Daniell cell produces electricity through	A. Electrolysis B. Radioactivity C. Spontaneous redox reaction D. Endothermic reaction
14	Standard electrode potential is measured under	A. 100 oC and 1 atm B. 0 oC and 1 M concentration C. 25 oC and 1 M concentration

		D. 100 oC and 1 M concentration
15	In a galvanic cell, the anode is	A. Negaive and site of oxidation B. Postive and site of oxidation C. Negative and site of reduction D. Postive and site of reduction
16	The process of coating a metal with zinc is called.	A. Alloying B. Galvanization C. Electroefining D. Electrplating
17	Electroplating involves.	A. Electrolysis for coating B. Spontaneous oxidation C. Redox without electordes D. Thermolysis
18	One Faraday is equl to	A. 1 C B. 96500 C C. 1 J D. 96500 J
19	If 1 Faraday of elecricity is passed the mass deposited equals.	A. 1 gram equivalent B. 1 gram C. 1 mole D. 1 atm
20	In redox reactions, oxidizing agents are	A. reduced B. Oxidized C. Always metals D. Always gases