Lab Safety and Practical Skills

Sr	Questions	Answers Choice
1	Which of the followign pairs would require methly orinage as a a indicator	A. Strong acid and strong base B. Weak acid and sgtrong base C. Strong acid and weak base D. Weak acia and weak base
2	The chromyl chlorikle test is a specific confirmatory tst for	A. Chloride ions B. Bromide ion C. Iodide ions D. sulfate ions
3	On dry heating test for salt analyshat turens limewater milky suggests the presence of.s, the evolution of coloriess, odouriess gas ta	A. Carbonate ion B. Chloridei on C. Sulfate ion D. Nitrate ion
4	A white precipitate that dissolves in excess ammonia is likely to be	A. Zinc hydroxide B. Calcium cabonate C. Lead Iodide D. Barium sulphate
5	Before starting a titration , the burette should be	A. Headed B. Drived completely C. Washed with distiled water only D. Rinsed with the solutin it will contain
6	While takig able a reading with a burette, why it is always advisibel to read the lower meniscus for the colorless liquids and the upper meniscus for the coloured liquids.	A. Because it is more convenient B. Becuse colorless liquids have more surface tension than colored liquids C. Because of the parallax effect D. Beacuse lower menisucs does not exist for colored liquids
7	Which of the followign is a primary standard solution.	A. NaOH B. Na2CO3 C. KOH D. HCl
8	Eating an drinking in the lab is	A. Allowed during breaks B. Allowed if you'r creful C. Strictly proibiteed D. Allowed int he chemistry lab but not physics
9	The addition of HCl to a carbonate salt leads to.	A. Effervescene=ce B. Blue colour formation C. White precipitate D. No visible reaction
10	Why phenolphtahalein indicator is more apropriate to se durng the titrations which inole a strong acid and a strong base.	A. Because it is iteslf weakly cidic B. Because the pH at ath equivalence point as well as the pH over where the colour of phenolphthalein changes match each oterh C. Becauwe teh solution at the end of titrationis acidic D. Beacuse the solutin at the end of titratic is basic
11	Which of the following is the most suitable indicator for the titration of a strong acid and a storng base.	A. Phenolphthalein B. Methyl orrange C. Universal indicator D. Litmus
12	The end point of a titration is indicated by	A. Formation of a gas B. Change in temperatutere C. Colour change of the indicator D. Change in pressure
13	Which of the followig is not a component of titration	A. Pipette B. Burette C. Test Tube D. Conicla flask

14	Which cation gives a white gelatinous precipitate upon the addition of aueoeous ammonia.	A. Cr3+ B. Cr2+ C. Zn2+ D. Al3+
15	Which cation forms a white precipitate with NaOH that dissolves in excess NaOH	A. Mg2+ B. Zn2+ C. Ca2+ D. Ba2+
16	What is the the typical volume of solution a pipette delivers in titration	A. 25 ml B. 10 ml C. 50 ml D. 100 ml
17	Why is it important to usea piptte in titration	A. To hold the acid B. To measure a fixed volume of the solution accurately C. To detect the endpoint D. To mix the acid and alkali
18	addition of NH4OH to an aqueous solution of cation gives a green precipitate which turns brown upon standing, Which basic radical is indicated.	A. Cr3+ B. Cu2+ C. Fe2+ D. Fe3+
19	Lab coats are important becausse they.	A. Are fashionable B. Prevent contamination and protect clothing C. Make you feel like a scientist D. Keep you warm
20	A gas with a pungent smell turns moist red litmus paperblue. The gas is likely	A. NH3 B. HCl C. CO2 D. SO2