Atomic Structure

Sr	Questions	Answers Choice
1	The shape of d-orbitals are like	A. Complicated B. Polar C. Spherical D. Cloverleaf
2	The quantum number 'm' of a free gascous atom is associated with	A. The effective volume of the orbital B. The shape of the orbital C. The spatial orientation of the orbital D. The energy of the orbital in the absence of a magnetic field
3	What information does the pricnipal quantum numerb (n) gives us about orbitals	A. size B. Shape C. Spin D. Size and shape
4	How many subshells are present in M-Shell	A. 1 B. 2 C. 3 D. 4
5	If we remove electrons from an atom until the nucleus is left, then it is called.	A. Ionization energy B. Ionization potential C. Successive ionization energy D. All
6	d-block elements are also called.	A. Inner transition B. Outer transition C. Typical transition D. None
7	Quantum numebr values for 2 p orbitals are.	A. n = 2, , l = 1 B. n = 1, 1= 2 C. n = 1, 1= 0 D. n= 2,1 = 0
8	How many d-orbitals are showing their lobes along axis.	A. 1 B. 3 C. 2 D. All
9	No two electrons in an atom have same all four quantum numbers, this is called.	A. Aufbau Principle B. Hund's rule C. Pauli's exclusing principle D. None
10	Orbital having same energy are called.	A. Hybrid orbials B. Valance orbitals C. Degenerate orbitals D. d-orbital
11	The value for (n) Principal quantum numebr can not be.	A. 0 B. 1 C. 2 D. 3
12	The value of 'm" cannot be	A. Fractional B. Zero C. Positive D. Negative
13	Two electrons in the same orbital have.	A. Same Spin B. Differece spin C. May same or different D. Parallel spin
14	1st ionization energy is.	A. Exothermic B. Endothermic C. Both D. Depend upon atom

15	Electronic cloud around teh nucleus is uniformly distributed in all directions in.....subshell	A. f B. s C. d D. p
16	How many d orbitals are there in a givne energy level.	A. 1 B. 5 C. 3 D. 7
17	s-subshell has.....orbital	A. 1 B. 2 C. 3 D. 4
18	Number of neutrons present in Cl are	A. 18 B. 17 C. 19 D. 25
19	Which of the following atoms show anomalous configuration.	A. Cu B. CO C. Fe D. Ni
20	The three p orbitals in a given energy level are oriented.	B. Along the same axis C. At 45^o each other D. Mutually perpendicular to each other along the x, y , and z axes. E. In a complex tetrahedral arrangement