States and Phases of Matter

Sr	Questions	Answers Choice
1	What property of wate allows quals life to survive under frozen surface	A. High surface tension B. High heat capacity C. Ice floats and insulates water before D. Low boiling point
2	What happens to the particles in a solid when the temperature increases.	A. They move closer to each other B. Their vibrational motion increases C. They start moving freely D. The solid melts
3	The shape in which a crystal usually grows is called its.	A. Habit B. Crystal lattice C. Melting Point D. Cleavage plane
4	At equilibrium in a closed system, which two processes occur at tha same rate.	A. Mellting and Freezing B. Evporation and condensation C. Sublimation and condensation D. Evaporation and boiling
5	Liquid crystals exhibit properties.	A. Only like solids B. Only like liquid C. Between solids and liquids D. Unliek solids or liquids
6	Which of the following statemetns is true about diffusion in liquids?	A. Liquids do not diffuse at all B. Liqudis diffuse faster than gases C. Liquids diffuse slower than gases due tos tonger intermolecular forces D. Liquids diffuse only when heated
7	Why can a water molecule formtwo hydrogen bonds on avarage.	A. Beacuse it is polar B. Due to its temperature C. It has 2 hydrogen atm and 2lone pairs D. Becuse it has a covalent bond
8	In Solids the particles miainly have which type of motion	A. Translatory motion B. Vibratory motion C. Rotatray motion D. Random Motion
9	Which of the following statemetns about ideal gases is true.	A. they have strong intermolecular forces B. Their particles have significant volume C. Their volume is mainly due to particle size D. They have negligible intermolecular forces
10	Why is water's surface tension so high.	A. Due to its ioic nature B. Because of hydrogen boding pulling surface molecules downward C. Beacue it has low boiling point D. Beacuse water molecules are very large
11	London dispersion forces are teh only forces present among.	A. Molecules of water in liquid state B. Atoms of helium in gaseous stat at high temperature C. Molecuels of solid iodine

		D. Molecules of hydrogen chloride gas
12	Surface tensin of liquid is due to.	A. Inword pull of surface molecule B. Upard pull from the surface C. Collision of molecules D. Repulsive forces
13	The expansion of liquids increase with temperatur eis.	A. Equa to that of gases B. Greater than that of gases C. Negligible copared to gasses D. Same as solids
14	Why does ice float on water.	A. It has lower density than water B. It is heavier than water C. It forms ionic bonds D. It denser than water
15	Liquid are approximately how many times less compressible than gases.	A. 10 B. 105 C. 100 D. 1000
16	Which state of matter exists only within a narrow range of temperature and presssure.	A. Liquid B. Solid C. Gas D. Plasma
17	Which of the following is a polar moleuce taht shows permanent dipole interactions.	A. CHCl B. CH4 C. Cl2 D. H2
18	What type of structue do water molecule sform in ice.	A. Linear B. Hexagonal close packed C. Regualr tetrahedral D. Irregualr amorphous
19	Why can gases be compressed easily.	A. Because they have strong intermolecular forces B. Becasue their particles are lightly packed C. Because most of their volume is empty space D. Beacuse they are denser than solids
20	Why are cholestric liuquid crystals useful in medical diagonostics.	A. They kill bacteria B. They reat with tumors chemically C. They detect temperatur edifferences D. They absorb X-rays