Chemical Equilibrium

Sr	Questions	Answers Choice
1	Law of mass action is applicable to.	A. Reversible reaction B. Combustion reactions C. Irreversible reactions D. Endothermic reactions only
2	Catalyst affects	A. Activation energy B. Equilibrium position C. Kp value D. Enthalpy
3	Which does not alter equilibrium constant.	A. Catalyst B. Temperture C. Pressure D. Both a and c
4	Na2CO3 in wate gives.	A. Acidic solution B. Basic Solution C. Neutral solution D. Buffer
5	Addition of inert gas at constant volume	A. Affects equilibrium B. Shifts reation left C. Shifts reaction right D. No effect
6	Combustion of methane is.	A. Reversible B. Endothermic C. Irreversible D. Equilibrium process
7	NH4Cl in water makes solution	A. Neutral B. Acidic C. Basic D. amphoteric
8	Catalyst affects	A. Value of K B. Equilibrium position C. Activation energy D. Final concentrations
9	Kc = 0.040 at 450 oC for the teh given reaction, evaluate Kp for the reaction. Pcl5 ..... PCl2 + Cl2	A. 0.40 B. 2.4 C. 0.64 D. 0.052
10	At equilirbium the observabel properties.	A. Keep changing B. Fluctuae randomly C. Remain constant D. Oscillate
11	Optimum temperatur ein Haber process is	A. 50 oC B. 450 oC C. 200 oC D. 1000 oC
12	Kp is used for.	A. Solid state reactions B. Gaseous reactions C. Aqueous reactions D. Liquid reactions
13	If Q=K , then	A. Reaction is irrversible B. System is at equilibium C. Forward reaction dominates D. Reverse reaction dominates
14	Contact process is used for	A. Sulfuric acid production B. Ammonia Synthesis C. Nitric Acid production D. Hydrogenation
15	Whcih of the following reactions reaches equilibrium	A. Reversible B. Irreversible C. Combustion D. Neutralization

16	Q < K implies.	A. Reaction proceeds forward B. quilibrium is established C. System stops D. Reaction fhifts in reverse
17	Catalyst used in Haber process is	A. Fe B. Pt C. Cu D. Zn
18	Which of the following statemetns correctly describes the effect of temperature on the equilirium constant.	A. Kc is directly proportional to temperature B. Kc is inversely proportional to temperature. C. Kc depends on the enthalpy change of the reaction D. Temperature has no effect on the value of Kc
19	Untis of Kc depend on	A. Catalyst B. Reaction stoichimometty C. Activation energy D. Delta H
20	Law of mass action applies to	A. Closed systems only B. Open systems C. Irreversibel systems D. Combustion reactions