Acid-Base Chemistry

Sr	Questions	Answers Choice
1	The pH of distilled water at 25 oC is	A. 6 B. 7 C. 5 D. 4
2	Which is a weak acid	A. HNO3 B. H2SO4 C. HF D. HCl
3	The solubility prodct of Pb SO4 is	A. 1.6 x 10⁻⁸ B. 1.4 x 10⁻⁴ C. 1.6 x 10⁻⁶ D. 1.6 x 10⁻²
4	A salt that hydrolyzes in water in from	A. Strong acid and base B. Weak acid or base C. None D. Both strong
5	Which of the following is the conjugate base of water.	A. OH- B. OH+ C. H2O D. H2O+
6	The solubility product of AgCl is 2.0 x 10^{-10} mol2 dm-6. The maximum concentration of Ag+ ions in the solution is.	A. 2.0 x 10⁻¹⁰ mol dm⁻³ B. 1.41 x 10⁻⁵ mol dm⁻³ C. 1.0 x 10⁻¹⁰ mol dm⁻³ D. 4.0 x 10⁻²⁰ mol dm⁻³
7	Acid rain is mainly due to	A. SO3 and NO3 B. CO2 C. H2 D. CH4
8	Which has the highest pH	A. 0.01 M NaOH B. 0.1 M NaOH C. 0.1 M HCl D. 0.01 M HCl
9	NaOH is not considered an Bronsted Lowry acid because.	A. It's neutral B. It's not soluble C. Doesn't donate H+ D. Doesn't produce H+
10	If the concentration of Cl- ion in a solution is increaed, the soubility of silver chloride will	A. Decreasse B. Increase C. Remain unchanged D. Become zero
11	pH of blood is around.	A. 6.0 B. 7.4 C. 5.4 D. 8.0
12	Litmus turns which color in a basic solution.	A. Red B. Orange C. Blue D. Colorless
13	pH of 1 x 10 -4 M HCl is	A. 4 B. 1 C. 3.5 D. 7
14	The conjugate acid of NH3 is	A. NH4 B. NH2 C. NO3 D. N2H4

15	The pH scale ranges typically from	A. 1 -10 B. 0 - 14 C. -1 to 1 D. 7 - 14
16	The Bronsted Lowry definition identifies acid as.	A. Electron pair acceptors B. Proton donors C. Proton acceptors D. Elecron pair donors
17	Which acid is present in vinegar	A. CH3COOH B. HNO3 C. H2SO4 D. HCl
18	The sgtrength of base is measured by its.	A. Ka B. Kb C. Pkw D. Kw
19	Which one is a Lewis base.	A. H+ B. NH3 C. HCl D. BF3
20	The conjugate base of HClO4 is	A. ClO4 B. ClO4 ²⁻ C. hClO3 D. H2ClO4+