MDCAT Chemistry Chapter 11 S and P Block Elements Online Test

Sr	Questions	Answers Choice
1	Oxidation state of Mn' in KMnO4. K2MnO4, MnO2 and MnSO4 is in the order	A. +7.+6.+2,+4 B. +6,+7,+2,+4 C. +7. +6.+4.+2 D. +4, +6, +7,+2
2	Paramagnetic behaviour is caused by the presence of	A. Unpaired electrons B. Paired electrons C. Paired protons D. Paired electrons in an atom. molecule or ion
3	Group of element belongs to IIB group	A. Zn. Cd. Hg B. Cu. Ag. Au C. Sc. Y. La D. Ni. Pd. Pt
4	When light is exposed to transtion element, then electrons jump from lower orbitals to higher orbitals in	A. Orbitals of s-subshell B. Orbitals of d-subshell C. Orbitals of p-subshell D. between different shells
5	Number of electrons involved in d-d transition of [Ti(H2O)6]+3	A. 1 B. 3 C. 2 D. 4
6	Zine does not show variable oxidation state, because	A. Its d-subshell is incomplete B. Its d-subshell is complete C. It is relatively soft metal D. It has two electrons in outermost shell
7	Which of these has at least one d electron	A. Sc+3 B. Mn+7 C. Ti+4 D. Cr+3
8	which one pair has the same oxidation state of-Fe?	A. FeSO4 and FeCl4 B. FeCl4and FeCl3 C. FeSO4 and FeCl2 D. Fe2(SO4)3 and FeSO4
9	d-d transition cannot be observed in	A. Cr B. Cu C. Mn D. Zn
10	Catalyst used for ammonia synthesis is	A. Cu B. Co C. Zn D. Fe
11	Transition compounds which occur as tripositive ions have no	A. 4s-electron B. 3p-electron C. 3s-electron D. 2s-electron
12	The number of unpaired electrons present in Fe ions is	A. 1 B. 2 C. 5 D. 0
13	[Ti (H20)6]3+ ion isin colour.	A. Yellow B. Blue C. Violet D. Red
14	In [Ti (H2O)]3+ which colour is transmitted	A. Yellow B. Blue and red C. Blue and yellow D. red and yellow
15	Stability of Cu-metal is due to filled of d-orbital	A. Half B. Completely C. Partially

		D. Quarterly
16	The strength of binding energy of transition elements depends upon	A. Number of electron pairs B. Number of unpaired electrons C. Number of neutrons D. Number of protons
17	Variable Oxidation state of is related to transition elements	A. empty d-subshells B. Completely filled C. Partially filled d-subshell D. d-d transition
18	Which of the following pair has the same no. of electrons in d- subshell	A. Sc+3,Ti+3 B. Mn+2,Fe+3 C. Ti+3,V+3 D. Cr+3.Co+2
19	No of unpaired electrons are maximum in	A. V+3 B. Mn+2 C. Fe+3 D. Cr+3
20	At which oxidation state Cu achieves electronic configuration of Zn+2	A. 0 B. +2 C. +1 D. +3