NAT I Engineering

Sr	Questions	Answers Choice
1	Which of the following value ofΔHf represent that the product is least stable?	A. -94.0 kcal mol⁻¹ B. -231.6 kcal mol⁻¹ C. +21.4 kcal mol⁻¹ D. +64.8 kcal mol⁻¹
2	If a salt bridge is removed between the two half cells, the voltage.	A. Drops to zero B. Does not change C. Increases gradually D. Increases repidly
3	Mild oxidation of glycerol with H2O2/FeSO4 gives	A. Glyceraldehyde B. Dihydroxy acetone C. Glycerose D. None
4	Portland cement is manufactured by using.	A. Limestone, clay and sand B. Limestone, gypsum and sand C. Limestone, gypsum and alumina D. Limestone, clay and gypsum
5	Calcium acetate when dry distilled given.	A. Formaldehyde B. Acetaldehyde C. Acetone D. Acetic anhydride
6	1-Chloropropane has two isomers, it is an example of	A. Functional group B. Position isomerism C. isomerismMetamerism D. Chain isomerism
7	Iron, once dipped in concentrated H2SO4 , does not displace copper from copper sulphates solution, because.	A. <span style="font-size:11.0pt;line-height:107%; font-family:"Calibri",sans-serif;mso-ascii-theme-font:minor-latin;mso-fareast-font-family: Calibri;mso-fareast-theme-font:minor-latin;mso-hansi-theme-font:minor-latin; mso-bidi-font-family:Arial;mso-bidi-theme-font:minor-bidi;mso-ansi-language: EN-US;mso-fareast-language:EN-US;mso-bidi-language:ER">It is less reactive than copper B. A layer of sulphates is deposited on it<o:p></o:p> C. An inert layer of iron oxide is deposited on it<o:p></o:p> D. All valence electrons of iron are consumed
8	Which one of halogen acid is a liquid/	A. HF B. HCl C. HBr D. HI
9	Evaporation of water is	A. An exothermic change B. An endothermic change C. A process where no heat changes occur D. A process accompanied by chemical reaction
10	The following has zero valency	A. Na B. Be C. Al D. Kr
11	Which of the following halogens does not form its oxyacids?	A. Fluorine B. Chlorine C. Bromine D. Iodine

12	Potassium crystallizes with a	A. Orthogonal lattice B. Cubic lattice C. Triclinic D. Ortho rhombic lattice
13	Phosphide ion has the electronic structure similar to that of	A. Nitride ion B. Fluoride ion C. Sodium ion D. Chloride ion
14	An endothermic reaction is one in which	A. Heat is converted into electricity B. Heat is absorbed C. Heat is evolved D. Heat is converted into mechanical work.
15	$\Delta H_{Naturalizationn}$is always	A. Positive B. Negative C. Zero D. Positive and Negative
16	The isomers due to the unequal distribution of carbon atoms on either side of the functional group belonging to the same homologous series are called.	A. Functional isomers B. Position Isomers C. Chain Isomers D. Metamers
17	The reaction/method that does not give an alkane is.	A. <span style="font-size:11.0pt;line-height:107%; font-family:"Calibri",sans-serif;mso-ascii-theme-font:minor-latin;mso-fareast-font-family: Calibri;mso-fareast-theme-font:minor-latin;mso-hansi-theme-font:minor-latin; mso-bidi-font-family:Arial;mso-bidi-theme-font:minor-bidi;mso-ansi-language: EN-US;mso-fareast-language:EN-US;mso-bidi-language:ER">Catalytic hydrogenation of alkanes B. Wurtz reaction <o:p></o:p> C. Hydrolysis of alkyl magnesium bromide D. Dehydrohalogenation of an alkyl halide <o:p></o:p>
18	Which of the following has greatest tendency to lose electron.	A. F B. Fr C. S D. Be
19	A current of 9.65 ampere flowing for 10 minutes deposits 3.0 g of the metal which is monovalent.The atomic mass of the metal is	A. 10 B. 50 C. 30 D. 96.5
20	Which of the following alcohols cannot be produced by treatment of aldehydydes or ketons with Nebh4 LiAH4	A. P2O2 B. H2SO4 C. CH3COONa D. CH2COOCH5