Chemistry (New Book) - 9th Class Chemistry English Medium Short Question Preparations

Q1. Justify with example that solubility of a salt increases with increase in temperature.

Ans 1: when a salt like KNO3 is dissolved in water, heat is absorbed. It means heat is required to break the attractive forces between ions of solute. There fore, solubility of such aslt increases with increase of temperature.

Q2. Define neutralization reaction.

Ans 1: Both acids and alkalis are known to cancel the properties of each other when mixed together in equla amounts. The reaction is called neutraization reaction. A salt and water are formed as a result of thsi reaction e.g.
Hcl + NaOH ----------NaCl + H2O

Q3. Define molar mass.

Ans 1: The mass of one mole of a substance is called as molar mass. e.g. Molar mass of O-atom =16 g.

Q4. What do you know about strong acids.

Ans 1: The acids which are ionize completely in wate are called strong acids. Like hydrogen chloride, Sulphuric acid and nitric acid behave as very strong acids becasue they ionize completely in water.

Q5. How ions and isotopes are formed.

Ans 1: Changing the numbe rof electrons of an element forms ions, While changing the numebr of neutrons of an element forms isotopes.

Q6. Which down the reaction between sodium oxide and water.

Ans 1: Sodium oxide dissolves in water to give sodium hydroxide.
Na2O+H2O ----------2 NaOH

Q7. Write down the conditions for equilibrium.

Ans 1: The following physicla cidnitions for eqilibrium is.
1. The concentration of reactant on product remains unchanged.
ii. The temperature fo the system remains constant.
iii. The pressure or volume of the system remains constant.
iv. Reaction must happen in closed system.

Q8. What do you know about chemicla energy.

Ans 1: The energy stored in a molecule in which atoms are bonded to each other is called chemicla energy.

Q9. How many atoms does our body replace every year.

Ans 1: Every year our body replaces about 98% of its atoms.

Q10. How many times eesium bigger than Helium.

Ans 1: The largest atom cesium is approximately nine times bigger than the smallest atom helium

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