Chemistry (New Book) - 9th Class Chemistry English Medium Chapter 4 Preparation

Q1. What is the difference between a mole and Avogadro's number?

Ans 1: Mole:
The atomic mass, molecular amss or formula mass of substances expressed in gram is called mole
It is unit of measurement

Avogadro's Number:
The numebr of particles present in one mole of a substance is called Avogadro's number
It is the number of particles in a mole

Q2. What is chemicla equation.

Ans 1: The chemists have developed a very suitable way of represetnting a chemicla change in terms of symbols of elements and formulas of compound. Representing a chemcial change in this way is called a chemical equation.

Q3. What is the significnace of stoichiometry.

Ans 1: The composition of all the chemicla products we use in our lives, such as shampoos, perfumes, soaps and fertilizers are formed using stoichiometic calculations. Without stoichiometry the chemical industry does not exist.

Q4. The empirical cormula of a compound is CH2O. It molar mass is 180 gmole-1.Determine its molecular formula.

Ans 1: Ans

Q5. The empirical formula of compund is CH2O. Its molar mass is 60 g mol-1. Determine its molecular formu.a

Ans 1: Ans

Q6. Define exmpirical formula with an example

Ans 1: The type of formula where show the simplest whole number ratio of atoms present in a compund is called empirical formu.a e.g. glucose has simplest ration 1:2:1 of carbon, hydrogen and oxygen respectively. Hence its empirical formula is CH2O

Q7. Determine the molecular formula of n-propyl alcohol. Its structural formula is.

Ans 1: CH3-CH2-CH2-OH

Q8. Write down the chemical formula of barium nitride.

Ans 1: The chemicla formula of barium nitride is Ba3 N2

Q9. Define molar mass.

Ans 1: The mass of one mole of a substance is called as molar mass. e.g. Molar mass of O-atom =16 g.

Q10. Define the law of conservation of mass.

Ans 1: The matter can neither be creaed not destroyed during a chemical reaction.
