Chemistry (New Book) - 9th Class Chemistry English Medium Chapter 5 Preparation

Q1. Are energy diagrams useful.

Ans 1: Energy diagrams are useful becasue it gives us immediate answer about energy changes happening in a reaction.

Q2. Difference between Heat and Enthalpy

Ans 1: Heat: This form of energy is released when a bond is formed and absorbed when bond is broken
It is not essential part of system
Heat just enter and leave the object.
It is represented by 'q'
Its unit is Joule

Ans 2: Enthalpy:
Enthalpy or heat content, is defined as the totla amount of thermal energy stored in a compound.
It is essential prt of system
Enthalpy change is equal to heat evolved or absorbed, at constant pressure.
It is represented by 'H'
It uint is Kj/mol.

Q3. Who use the word energy for the first time.

Ans 1: Thomas Young was the first to use the word energy to the field of physics in 1802

Q4. Why the chemical reaction between sodium metal and water proceeds violently.

Ans 1: Reaction between sodium and water is exothermic that's why sodium react violently with water.

Q5. What is the role of glycogen in our body.

Ans 1: Glycogen is primary storage form of glucose. It is stored in liver and muscle. It help to regulate blood sugar level in our body.

Q6. What is chemicla energetic.

Ans 1: The study of energy chanes that take place during a chemicla reaction is called chemicla energetic.

Q7. Define Activation Energy.

Ans 1: Energy absorbed by reactant or product molecule in order to be converted into transition state is called activation energy. It is represented (Ea)

Q8. What is meant by enthalpy?

Ans 1: The total heat content of a body is called enthalpy of body. It is represented by H.

Q9. What is teh role of glycogen in our body.

Ans 1: The glycogen is the primary storae form of glucose. It serves as a source of carbohydrates n the body. It serves as an energy reservoir.

Q10. Why is breaking of a bond an endothermic process.

Ans 1: Bond breakign is an endothermic process because when two atoms combine, energy is stored in them and when we break bond, equal amount of energy is given to break that bond this energy is also known as dissociation energy. e.g. dissociation energy for hydrogne is 435 kJ/mol

---