Chemistry (New Book) - 9th Class Chemistry English Medium Chapter 7 Preparation

Q1. Why does ammonium hydroxide only partially ionize in water.

Ans 1: Ammonium hydroxide is only partially ionized in wate becasue it is a weak alkali.

Q2. How bases react with acid? Give two reactions.

Ans 1: Ans:

Q3. How can you justify that NH3 is Bronsted-Lowry base but not Arrhenius base.

Ans 1: Ammonia is Bronsted-Lowry base because it has the ability to accept a Proton but not Arrhenius base beacuse it does not produce hydroxide ion OH is aqueous solution.

Q4. Give two examples of Bronsted-Lowry bases which are not bases by Arrhenius definition.

Ans 1: Ammonia and Chloride ion are Bronsted Lowry bases which are not Arrhenius bases.

Q5. Write reaction of acids with carbonates and bicarbonates.

Ans 1: Ans

Q6. Write down the uses of oxalic acid.

Ans 1: Oxalic acid is the simplest organic diprotic acid. Its commerical uses include bleaching straw and leather and removing rust and ink stains from fabries.

Q7. Choose Arrhenius acids amont the following compunds HF,NH4,H2SO4,SO3,H2S,H2O

Ans 1: Arrhenius acids are substanes which dissociate in water to give proton. (H+)
Among the compunds, the Arrhenius acids ae.
1. Hf
2.H2SO3
3. H2S
4. H2O

Q8. Define Conjugate base

Ans 1: The conjugate base of strong acid is typical weak and does not readily accept protons, while the conjugate base of a weak acid is stronger and can accept protons more easily.

Q9. In what ways are mineral acid useful for us?

Ans 1: Minearl aciss are sued in several industrial processes.
In food industry mineral acids are used as preservative and as flavor agent.
As cleaning agen
In lead storage batteries.

Q10. What are bases according to Arrhenius concept.

Ans 1: Bases: The substances which can produce bydroxide ion (OH-) e.g. NaOH
